

Q-2 Answer any 5 Questions.

[4 mark each]

(20)

1. Present this value through pie graph :- Hexane 45 ml, Methanol 24 ml, Acetone 12 ml and Acetonitrile 9 ml.
2. 0.1 N $K_2Cr_2O_7$ in 500 ml is there, if we will dissolve it in more 500 ml then what will be the molarity of the newly formed solution.
3. A glass plummet weighs 14.59 g in air, 9.12 g when immersed in water, and 11.78 g when immersed in organic solvent. Calculate specific gravity of the organic solvent.
4. Define percentage dilution. How to make 3.68 % sucrose solution 750 ml water.
5. 5 ppm iron solution in water is present; how many % and how many molar the same solution is?
6. How to make 2.2 N H_2SO_4 ? And convert it into molar, ppm and ppb also.

Q-3 Answer any 4 Questions.

[7 mark each]

(28)

1. Discuss atomic and molecular weight & find out the molecular weight of $C_6H_{12}O_6$.
2. What is ppm concentration? If 10mg of NaCl dissolve in 2 ml water then how many ppm it will become and convert it into ppb also.
3. What are IV solutions? Discuss it through appropriate example.
4. What is specific gravity? Discuss it through proper example.
5. What is graphical presentation give its detail by the types of different graphs.

Atomic Weights:

H-1, C-12, O-16, Na-23, S-32, Cl-35.5, K-39, Cr-52.
